

Ph and poh calculations worksheet answer key class

I'm not robot!

KOH is an example of a strong base, which means it dissociates into its ions in aqueous solution. Although the pH of KOH or potassium hydroxide is extremely high (usually 10 to 13 in typical solutions), the exact value depends on the concentration of this strong base in the water. Therefore, it is important to know how to calculate the pH. What is the pH of a 0.05 m solution of potassium hydroxide? Potassium hydroxide or KOH, is a strong base and will dissociate completely in water at K+ and OH-. For every mole of KOH, there will be 1 mole of OH-, so the concentration of OH- will be the same as the concentration of KOH. Therefore, [OH-] = 0.05 M. Since the OH- concentration is known, the pOH value is more useful. POH is calculated by the formula $\text{pOH} = -\log[\text{OH}^-]$ Enter the concentration found before $\text{pOH} = -\log(0.05)$ $\text{pOH} = -(-1.3)$ $\text{pOH} = 1.3$ The value for pH is required and the relationship between pH and pOH is given by $\text{pH} + \text{pOH} = 14$ $\text{pH} = 14 - \text{pOH}$ $\text{pH} = 14 - 1.3$ $\text{pH} = 12.7$ The pH of a 0.05 m solution of potassium hydroxide is 12.7. The pH is a measure of how acidic or basic a chemical solution is. The pH scale ranges from 0 to 14. A value of seven is considered neutral, less than seven acids and greater than seven base. The pH is the basic 10 negative logarithm ("record" on a calculator) of the hydrogen ion concentration of a solution. To calculate it, take the log of a given concentration of hydrogen ions and invert the sign. See more information on the pH formula below. Here is a more in-depth review of how to calculate pH and what pH means in relation to the concentration of hydrogen ions, acids and bases. There are several ways to define acids and bases, but pH refers specifically only to the concentration of hydrogen ions and is applied to aqueous (water-based) solutions. When water dissociates, it produces a hydrogen ion and a hydroxide. See this chemical equation below. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ When calculating the pH, remember that [] refers to the molarity, the molarity M, is expressed in units of di usually per litre of solution. If you are given the concentration in a unit other than moles (mass percentage, molality, etc.), convert it to molarity to use the pH formula. The relationship between pH and molarity can be expressed as: $K_w = [\text{H}^+][\text{OH}^-] = 1 \times 10^{-14}$ at 25°C for pure water $[\text{H}^+] = [\text{OH}^-] = 1 \times 10^{-7}$ The equilibrium equation gives the following formula for pH: $\text{pH} = -\log_{10}[\text{H}^+]$ $[\text{H}^+] = 10^{-\text{pH}}$ of the molar concentration of hydrogen ion or the molar concentration of hydrogen ion is equal to 10 to the power of the negative pH. It's easy to do this calculation on any scientific calculator because most of the time, they have a "log" button. This is not the same as the "ln" button, which refers to the natural logarithm. You can easily use a pH value to calculate the pOH if you remember: $\text{pH} + \text{pOH} = 14$ This is especially useful if you are asked to find the pH of a base, since you usually solve the pOH rather than the pH. Try these problems to test your knowledge of pH. Calculate the pH for a specific [H+]. Calculate the given pH [H+] = 1.4 x 10-5 M Answer: pH = -log10[H+] = -log10 (1.4 x 10-5) pH = 4.85 Calculate [H+] from a known pH. Find [H+] if pH = 8.5 Answer: [H+] = 10-pH [H+] = 10-8.5 [H+] = 3.2 x 10-9 M Find the pH if the concentration of H+ is 0.0001 moles per liter. Here it helps to rewrite the concentration as 1.0 x 10-4 M because that makes the formula: pH = -log10(1.0 x 10-4) = 4. Or, you could just use a calculator to take the log. This gives you: Answer: pH = -log10(0.0001) = 4 Usually, you are not given the concentration of hydrogen ions in a problem, but you have to find it from a chemical reaction or from an acid concentration. The simplicity of this will depend on whether you have a strong acid or a weak acid. Most of the problems that require pH are about strong acids because they dissociate completely into their ions in the water. The weak acids, on the other hand, only partially, so equilibrium solution contains both the weak acid and its ions. 7-8262-285-0, 7-8262-285-0 NBSI, llah Ectitnerp, Kroy Yes,)De H16(Sisylna Lacimehc Evitatitnaug s'legov, k, k, Samoht, ,samoh; : J, sentab; Erup Fo Noinu Lanotaretini1350555891Cap/1531.01:iod, 245ââ135 :)(75 .mehc .lppa erup .ygonimret detailed Dna ,hp s. A. R. tsrud; Seulav hp ,hp Evitagen Etuceteroeh ylloceroeht s'ti elhw)31 ot 11 Dnuora yllaus(Eucal Hp HGIH A EVAH DNA DNA jeerht hcum hcum. Srewna ruoy erus ekam swawla ,snoitaluclac hp gnimrofrep e R'uoy neh 5.1 = hp)30.0(Gol - = hp m 30.0 =) +h[:rewna .Notito dica noitatenehc eht sa emas emas emas yltcaxe snoi edryh , . Ralom 1:1 ot Gnidrocca setaicossid taht dica gnorts that si dica cirolhcordyh ,rebmemer ,dica cirolhcordyh Fo nouloh m 30.0 a FO Hp Eht dnif .setaicossid of the

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